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| Review Unit 3-Chapter 10,11.1,11.2,12.1 Chem 11- Mrs. Sanford Name: |
| Mole- |
| Avogadro’s number- |
| Representative particle means atoms, ions, molecules, formula units, and so on. |
| Conversions:Fill in mol |
| Formulas for conversions involving moles🡪 n=\_\_\_\_ n=\_\_\_\_\_\_  M 6.02 x 10 23 |
| Note : section 10.2 was done in the gas unit |
| Percent composition: |
| Formula for % composition🡪 |
| Empirical Formula🡪 give the steps 1) % to mass 2) 3) 4) |
| Molecular formula can be drawn with a structural diagram, whereas an empirical formula cannot. |
| Molecular Formula- |
| In a chemical equation *reactants*🡪 ? Please review symbols at the bottom of p 323 |
| Catalyst- |
| Where is a catalyst written with an equation? |
| Coefficient- |
| The five types of reactions are(write below) and give an example for each. |
| 1) |
| 2) |
| 3) |
| 4) |
| 5) |
| Please go over page 338-339 Answer the following: |
| For reaction 1, give the another name for the **type** of reaction example that they give🡪 |
| For reaction 2, what is the liquid that is formed? |
| For reaction 3, what is another way to write water for predicting the products? |
| For reaction 4, what is the precipitate? |
| For reaction 5, what is burning in the bunsen burner? |
| What does mass conservation mean?(see p 357) |
| Using Fig 12.3 on p 357 Do #9 on p 358 below: |
| 🡪2K(s) + 2H2O(l)🡪 2KOH(aq) + H2(g) |
| #of particles 🡪 |
| Moles🡪 |
| Masses(need molar mass first)🡪 |
|  |
| Do p 306 #32, 33, p 307 # 35 p 310 #36,37, p 312 38,39, #45 (also draw a structural diagram of this molecule) Do p315 #52a, 53a,63c, #66(also draw a structural diagram of this molecule using molecular formula)p 316 #73(be careful with this one)p 347 Do 52, p 349 Do #71all |